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It would greatly facilitate the editors' task and thus be to the advantage of the author if, when submitting a paper which is part of a series, the author were to enclose reprints of the preceding immediately relevant parts.


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At the end of the paper references should be given in alphabetical order according to the name of the first author of the publication quoted, and should include the authors' initials, but not the title of the paper. Titles of journals should be abbreviated in accordance with the system used in the World List of Scientific Periodicals (1935, 2nd ed., Oxford: University Press). Examples of such abbreviations will be found in the current numbers of the Biochemical Journal, and a useful list has been issued by the Biological Council and is obtainable from the Editorial Office, Biochemical Journal for the National Institute for Medical Research, The Ridgeway, Mill Hill, London, N.W. 7, price 2s. 6d. post free. References to books and monographs should include the town of publication and the name of the publisher, as well as the date of publication and the number of the edition to which reference is made. Thus:


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Chemical Formulae. These should be written as far as possible on a single horizontal line. With inorganic substances and CHCl₃, CCl₄ and CS₂, formulae may be used in the text as abbreviations, particularly in the experimental portion, at the discretion of the editors. With salts it must be stated whether or not the anhydrous material is used, e.g. anhydrous CuSO₄, or which of the different crystalline forms is indicated, e.g.

CuSO₄, 5H₂O, CuSO₄·H₂O.

Moles. The word 'mole' which has not previously been used in the Biochemical Journal, will not be permitted with the meaning of 'gram-molecule'. It must not be abbreviated to 'mol'. The sub-multiples should be printed as 'm-mole' and 'µ-mole'.

The number of molecular proportions of reagents employed in a reaction, often referred to in everyday language as 'molecules', may be abbreviated in Experimental sections to 'mol-prop.'.

Description of Solutions. Solutions of common acids and bases should always be expressed in terms of normality (N), and salts preferably in terms of molarity (M), e.g. N-HCl; 0·1 m-NaH₂PO₄. Fractional concentrations should preferably be expressed in the decimal system, e.g. 0·25 N-HCl (not N/4 HCl). The term % must be used in its correct sense, i.e. g./100 g. of solution. For 'per cent by volume', i.e. ml./100 ml., the term % (v/v) may be employed. To indicate that a given weight of substance is contained in 100 ml. of solution, the term % (w/v) (weight per volume) may be used.

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The virulence-enhancing factor of mucins:
5. The different components of the 'third factor' involved in virulence enhancement. By H. SMITH, R. C. GALLOP and J. L. STANLEY
6. Identification of a heparin as the main component of the third factor involved in virulence enhancement. By H. SMITH, R. C. GALLOP, PATRICIA W. HARRIS-SMITH and J. L. STANLEY

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